

# Electron Configuration

**Electron Configuration:** The arrangement or distribution of electrons across orbitals or energy levels around the nucleus is known as electron configuration.

The electron shells are labeled K, L, M, N, O, P and Q. The four different types of orbitals (s, p, d, and f) have different shapes, and one orbital can hold a maximum of two electrons. The p, d, and f orbitals have different sublevels, thus can hold more electrons.

- \* 1st shell has one orbital
- \* 2nd shell has s and p orbitals
- \* 3rd shell has s, p and d orbitals

The pattern is that the number of orbitals in each shell is equal to the square of the shell number.

- \* 4th shell has s, p, d and f orbitals
- \* 5th shell has s, p, d and f orbitals
- \* 6th shell has s, p, d and f orbitals
- \* 7th shell has s, p, d and f orbitals

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## (Aufbau) Principle

energy levels

